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Extraction of limonene from orange peel by steam distillation lab report. Steam distillation formula. Example of steam distillation. Isolation of limonene by steam distillation lab report.

Full PDF PackageDownload Full PDF PackageThis PaperA short summary of this paper37 Full PDFs related to this paperDownloadPDF Pack In this lab, extractions, classification tests, and a distillation was performed in order to isolate two natural products, eugenol and limonene. In the isolation of limonene, steam distillation was used to obtain the water insoluble oil. This was done with 10.01g of orange peel and 50mL of water. Once heated, a 30mL distillate was obtained. This was a mixture of water and the organic compound of limonene. In order to isolate the limonene, 5.0mL of ethyl acetate was added to the water and limonene mixture and two layers were separated. The two layers that formed were clear in color, but had distinct consistencies. The bottom layer looked very thick, while the top looked thin and moved quicker than the bottom layer. The top layer is to one that...show more content...To obtain these results, Eugenol was isolated from a mixture of clove oil and dichloromethane. This was done by extracting the 20mL mixture twice with 15mL sodium hydroxide both times. At first the solution was slightly tan in color. After the addition of NaOH, a yellow and clear transparent layer formed. The clear layer was thrown out, and the yellow layer was then mixed with 15mL oh HCl. This solution created a white and brown layer with in the flask. Once Dichloromethane was added, the solution lost the brownish color and created a solution with a white, slightly transparent layer and a yellow layer. Only the yellow layer was kept and after being heated, it was marigold in color. This solution was heated to remove what was left of the dichloromethane, and resulted in pure Eugenol. The liquid smelled very spicy and rich. In this lab, the final mass of the Eugenol was not calculated because therefore the percent Eugenol in clove oil could not be calculated. It order to calculate the percent of eugenol in clove oil, one should take the mass of eugenol in grams and divide it by the original 2g of clove oil and then multiply by 100. To prevent this from happening again, one should pay attention to the instructions in the lab manual and not skip steps. Once this error was realized, it was too late because all of the test had been done on the Eugenol and so a lot of what recovered was...show more content...The Bromide Test and permanganate test were performed with cyclohexane and cyclohexene, to further understand how these tests work. Both test were positive for the cyclohexene and negative for the cyclohexane. These test were positive for the Cyclohexene because of the Double bond present. This double bond, which is present in limonene and eugenol as well, reacts with the bromine and permanganate to produce a new compound. The production of the new compound id confirmed by the fading of the brown color of bromide and the change from purple to brown of the permanganate. These test did not work for the cyclohexane because no double bond was present. The lack of the double bond means that the reaction had not place to occur and therefor did not. These standards help in the understanding of what the outcomes of Eugenol and Limonene should be. Since these outcomes were reasonable and both compounds were purified, this lab was Natural Products and AlkenesFelicia George O-chem 237- Red 11/7/17 For this lab is to isolate examples of two types of natural products terpenes and acetogenic. The properties of these two natural products will also be studied. For this experiment to happen a steam distillation and extraction of a weak organic acid will be used to isolate two natural products (limarene & eugenol) classification test will be used to provide the evidence for some functionality present in these structures. Limonene has a pleasant scent, its water-insoluble oil that's found in many plants (Organic Chemistry Experiment 2017). Limonene is spotted in concentrations in the range to 9700ppm which is 0.97% by weight in peels of citrus fruits (Organic Chemistry Experiment2017). For this lab Limonene was obtained by the steam distillation of orange peels and dichloromethane extraction of the orange peel. For eugenol a spicy- smelling phenol which is the main constituent of oil of cloves (Organic Chemistry Experiment 2017). Eugenol is widely used as dental analgesic, that's the smell of the dentist office. For this lab, eugenol will be obtained by a base extraction of clove oil. Isolation of Eugenol: To start the experiment dichloromethane was extracted from 20ml of dichloromethane of clove oil by using two 15ml portions of 3M NaOH. Water was added to make the layers separate from each other. Dichloromethane was at the bottom and the aqueous on top. Then the aqueous layers were combined and 6M

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